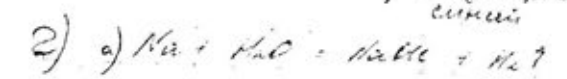
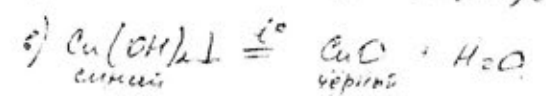
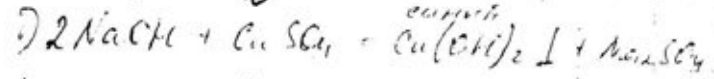


Российская Федерация  
 Министерство образования  
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 "\_\_\_" \_\_\_\_\_ 20\_\_ г.  
 N" \_\_\_\_\_

№10-1

ТсБ-Х10-27-34

- 1) А - Na (натрий) 0,5  
 Б - Cu (медь) 1,5

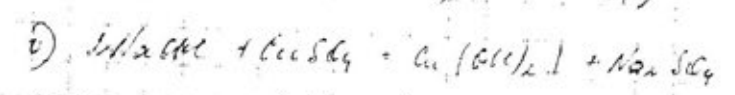


1) n(Na) =  $\frac{m}{M} = \frac{9,2}{23} = 0,4$  (моль) изоб.

2) n(H<sub>2</sub>) =  $\frac{V}{V_{н.у.}} = \frac{4,48}{22,4} = 0,2$  (моль) изоб.

3) n(NaOH) = n(H<sub>2</sub>) = 0,2 (моль)

n(NaOH) = n(Na) = 0,2 моль = 4,6 г



1) n(Na<sub>2</sub>SO<sub>4</sub>) =  $\frac{n(NaOH)}{2} = \frac{0,2}{2} = 0,1$  (моль) = n(Cu(OH)<sub>2</sub>)

2) m(Cu(OH)<sub>2</sub>) = n · M = 0,1 · 98 = 9,8 г

3) m(Na<sub>2</sub>SO<sub>4</sub>) = n · M = 0,1 · 142 = 14,2 г

4) m(Na<sub>2</sub>SO<sub>4</sub>) + m(Cu(OH)<sub>2</sub>) = 14,2 + 9,8 = 24 г

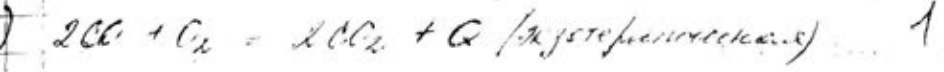
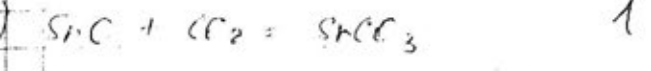
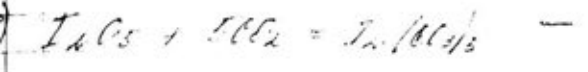
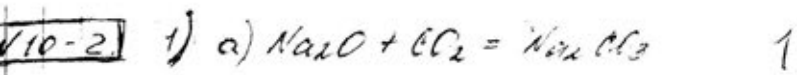
5) ω(Cu(OH)<sub>2</sub>) =  $\frac{m_{в.в.}}{m_{смеси}} \cdot 100\% = \frac{9,8}{24} \cdot 100\% = 40,8\%$

ω(Na<sub>2</sub>SO<sub>4</sub>) =  $\frac{m_{в.в.}}{m_{смеси}} \cdot 100\% = \frac{14,2}{24} \cdot 100\% = 59,2\%$

Субст. 1. А - Na, Б - Cu

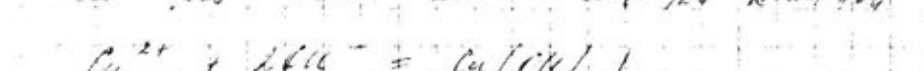
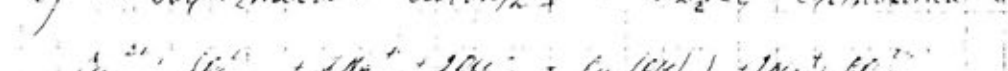
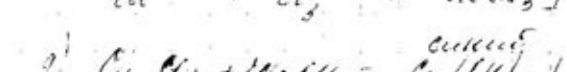
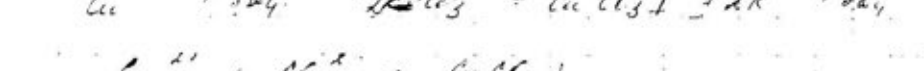
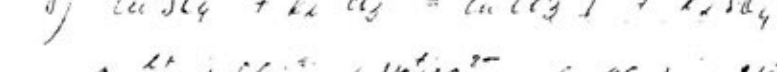
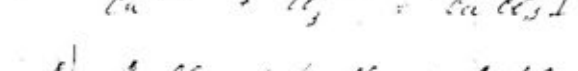
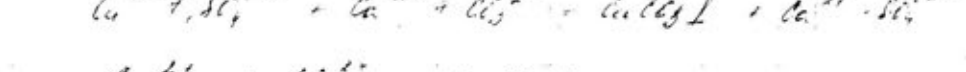
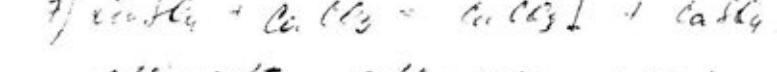
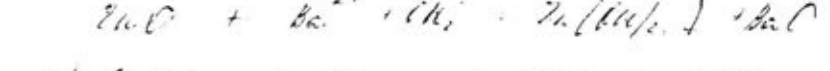
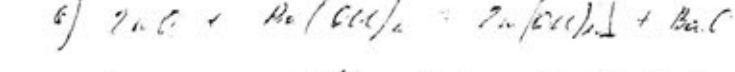
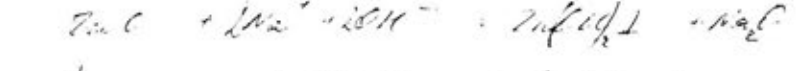
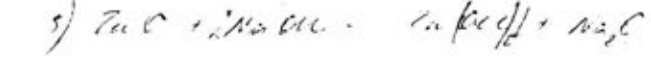
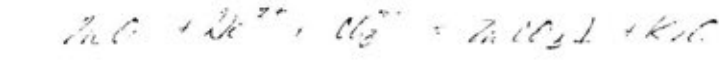
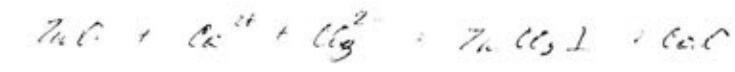
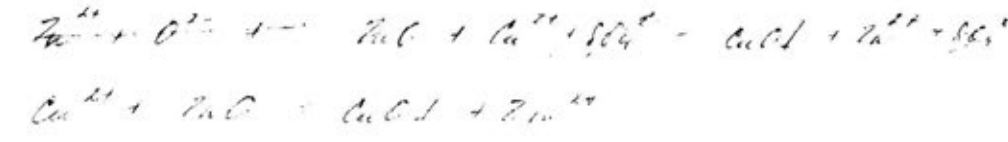
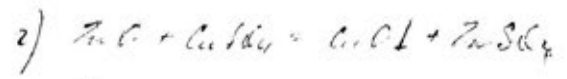
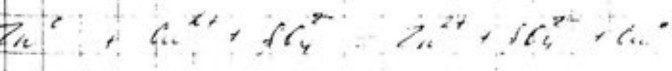
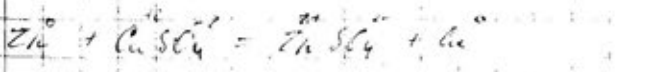
1	2	3	4	5
4	4	5	8	5

0,5  
 1,5  
 0,5  
 1,5  
 0,5  
 1,5



0-4) 1) уросты по легкости окисления:  $\text{CuSO}_4 > \text{Fe} > \text{Ba} > \text{Ba}(\text{OH})_2$  - дробить, Zn не окисляется

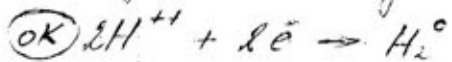
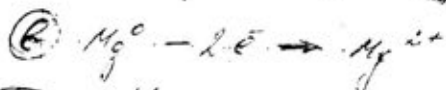
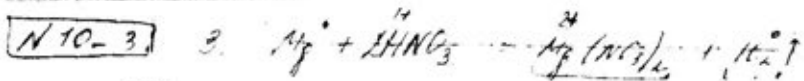
	Zn	ZnO	$\text{CuSO}_4$	$\text{CaCO}_3$	$\text{K}_2\text{CO}_3$	$\text{NaHCO}_3$	$\text{Na}_2\text{SO}_4$	$\text{NaOH}$	$\text{Ba}(\text{OH})_2$
Zn	-	-	+	-	-	-	-	-	-
ZnO	+	+	-	-	-	-	-	+	+
$\text{CuSO}_4$	-	-	-	-	-	-	-	-	-
$\text{CaCO}_3$	-	-	+	-	-	-	-	-	+
$\text{K}_2\text{CO}_3$	-	-	-	-	-	-	-	-	-
$\text{NaHCO}_3$	-	-	-	-	-	-	-	-	-
$\text{Na}_2\text{SO}_4$	-	-	-	-	-	-	-	-	+
$\text{NaOH}$	-	+	+	-	-	-	-	-	-
$\text{Ba}(\text{OH})_2$	-	+	+	-	-	+	-	-	-



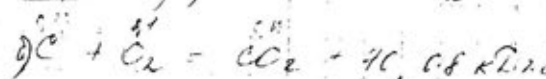
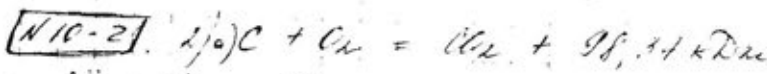
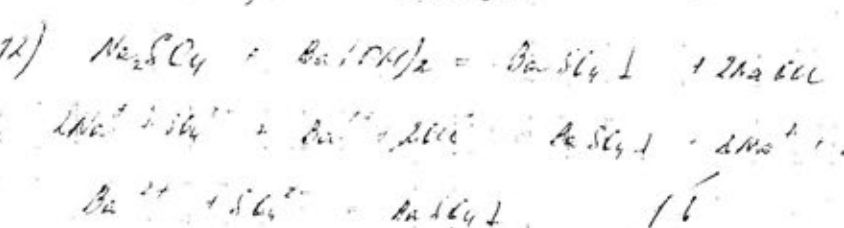
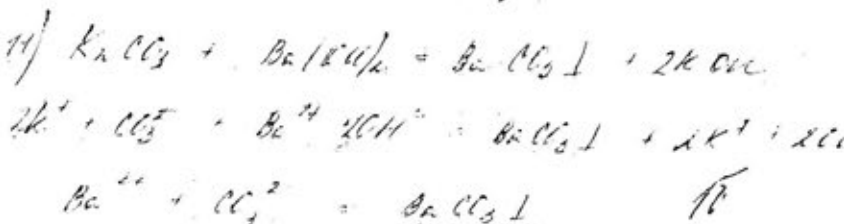
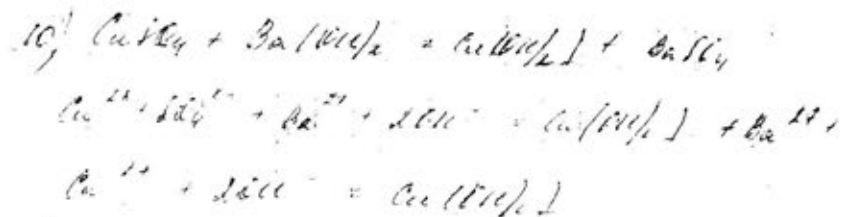
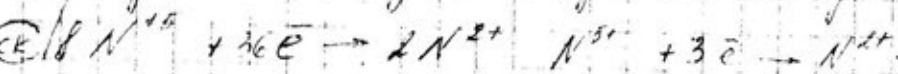
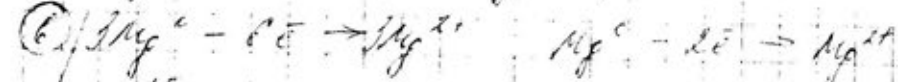
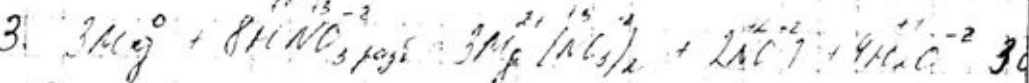
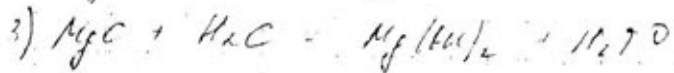
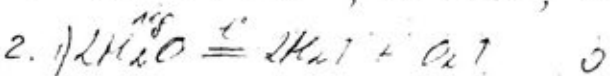
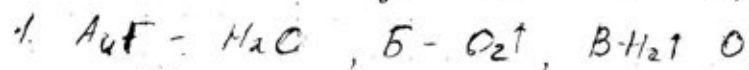
продолжение



4)  $Al_2(SO_4)_3$  - соль, растворима в воде, имеет кислую среду т.к. Al амфотерный, S-кислотный, с.миссия



Вещество Г -  ~~$2HNO_3$~~   $H_2$  - газ

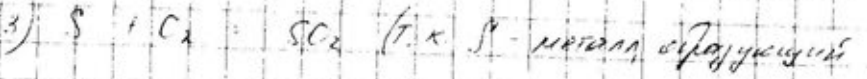
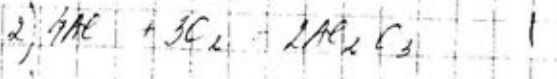
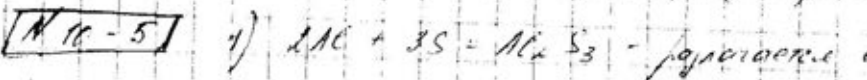


1.  $n(CaCO_3) = \frac{15}{100} = 0,15 \text{ (моль)}$

2.  $n(CaCO_3) = n(CO_2) = 0,15 \text{ (моль)}$

3)  $n(C) = \frac{3}{12} = 0,25 \text{ (моль)}$

4.  $n(Cu) = n(C) + n(H_2) = 0,25 + 0,15 = 0,4 \text{ (моль)}$



соединения, которые при взаимодействии с металлами образуют красные вещества) + высокая температура